

COMPOSITION OF COMET 67P/CHURYUMOV-GERASIMENKO REFRACTORY CRUST AS INFERRED FROM VIRTIS-M/ROSETTA SPECTRO-IMAGER

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Abstract

We report the interpretation of VIRTIS/ROSETTA spectra of comet 67P/CG, based on laboratory data collected on meteorites and series of analogs. We show that the crust is composed by a macromolecular carbonaceous material, mixed with minerals and likely more volatile molecules and possibly ammonium salts.

1. Introduction

The Visible InfraRed Thermal Imaging Spectrometer, VIRTIS [1] onboard ESA's Rosetta orbiter has completed two extensive mapping campaigns of the 67P/Churyumov-Gerasimenko comet nucleus in August-September 2014 [1,2]. Reflectance spectra were collected within the 0.4-5 μm range and reveal a dark surface (normal albedo $\sim 6\%$), a positive (red) near-infrared spectral slope, a steeper red visible slope, and a broad feature that peaks at 3.2 μm [1]. Here we address interpretation of these spectra through comparison with experimental data.

2. VIRTIS spectra

A full image-cube, with thermal contribution removed [3] from the MTP06 observing campaign (August 2014) has been analyzed in order to retrieve the general behavior of the surface spectra. The main spectral features of this data set are a single normalized infrared slope, two classes of normalized visible slopes ("neck" and "body") and a 3.2 μm band depth that appears to vary independently of the slopes for the body.

3. Organics and opaque minerals

Dark surfaces of small solar system bodies are often

related to the presence of dark refractory organics. Primitive chondrites contain a polyaromatic black solid (1-5 wt%), which is insoluble in common organic solvents and extracted with chemical HF/HCl protocols (IOM, Insoluble Organic Matter). IOM shares similarities with terrestrial type III kerogens and coals, and coal samples have been used as fair analogues in earlier studies [4]. Raman spectroscopic analyses clearly evidenced polyaromatics compounds are present in grains of plausible cometary origin (stratospheric IDPs, Antarctic micrometeorites), but likely with a broader range of chemical compositions than chondritic IOM [5]. Therefore, we selected a series of coal samples from the Penn State University Coal Bank and Data Base covering a wide range of maturities, i.e. a wide range of chemical compositions and degrees of structural order of the polyaromatic structures. The reflectance spectra of these coal samples were collected with the Spectrogonio Radiometer at IPAG at normal incidence and 30° emergence.

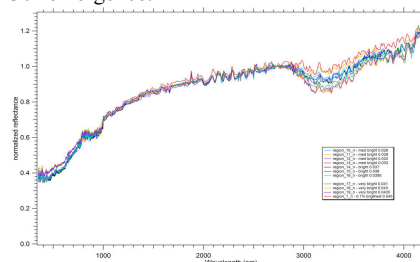


Fig. 1: representative nucleus surface vis+IR spectra ("neck" and "body") normalized at 2.72 μm with thermal emission removed. High frequency spectral variations are due to noise and calibration artefacts.

The reflectance spectra strongly depend on maturity and grain size [7,9]. All samples display very low reflectance in the visible range. In contrast, low maturity samples display a rather high

reflectance level in the near-infrared range, while mature coals (e.g., anthracite) are spectrally flat and dark over the visible and near-infrared ranges. Mature coals cannot be considered as fair analogues of cometary refractory organics. Mature IOM is recovered in thermally metamorphosed chondrites that experienced high temperatures (> 600 °C) conditions over millions of years. The near-infrared range appears then as a critical range for constraining the nature of the dark components of 67P/CG. Indeed, the reflectance spectra of primitive carbonaceous chondrites and of IOM extracted from Orgueil CI chondrite are basically flat and dark in the visible and near-infrared [6].

A strong effect of grain size is observed. The reflectance level in the near infrared is found to increase upon decreasing grain size, in particular in the case of a lignite whose composition is very close to that of the IOM of the CI Orgueil chondrite. As submicrometric particles are expected in the 67P/CG refractory crust, we infer that grain size cannot account for the high near infrared reflectance. Similar observations have been made on a wide range of terrestrial organics like solid oil bitumens, and that our general conclusions obtained from a series of coals would apply to polyaromatic solids that sample broader range of compositions and polyaromatic structures [7].

Furthermore, dark polyaromatic hydrocarbons that are ubiquitous in cometary and asteroid grains cannot account for the low reflectance of primitive chondrites, and of comet 67P/CG. In fact, the very low reflectance and spectrally flat reflectance spectra of primitive carbonaceous chondrites and of their IOM strongly suggests spectral contribution from opaque minerals such as sulfides and oxides, which cannot be dissolved using HF/HCl protocols. Sulfides (troilite, pyrrhotite, pentlandite) and Fe-Ni alloys (kamacite, taenite) are ubiquitous in chondrites and cometary grains. They play a key role in the control of the low albedo of carbon-rich cosmomaterials [8,9], and likely for comet 67P/CG. Finally, sulfides can also exhibit red slopes in the visible and might account for the spectral characteristics of the spectra of comet 67P/CG [1].

4. The 3.2 μm band

The broad 3.2 μm band can be assigned to OH, CH, H₂O, NH/NH₂ and NH₄⁺ chemical groups, molecules or ions. The contribution of water ice is considered as

weak due to the lack of peak at 3.1 μm . Water ice has been marginally detected in area of the nucleus just emerging from shadow [10]. NH/NH₂ do not fit properly the 3.2 μm band, and their ubiquitous presence is also unlikely, as N-rich cometary grains are extremely scarce in collections [11]. The OH chemical groups trapped in minerals or inserted within the network of refractory polyaromatic solids is a possible candidate. OH in proto-serpentine and glasses, which are stable up to 300 °C, are reported in primitive chondrites [12, 13]. However, the peak position of their band lies around 3.1 μm , and the width is smaller than that of the observed 3.2 μm band. In a carbonaceous structure, OH can be present as either alcohol (terminating -OH) or carboxylic groups (-COOH). The carboxylic band is broader than the alcoholic one, due to efficient proton exchange through hydrogen bond. The contribution of carboxylic groups appears thus highly plausible, but contribution of alcoholic OH cannot be rejected. Ions like NH₄⁺ could also contribute to the 3.2 μm band. This ion spectrally matches well the 3.2 μm band, and has been found to form along with OCN⁻ in laboratory experiments [14]. Tosum up, OH and alkyl groups in organics and ammonium salts appear as the most plausible candidates.

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